

Factors Affecting Reaction Rates Name_____

- 1) If two separate atoms are floating around in a container, when might they form a bond? What do they have to do?

- 2) List two ways that you could increase the concentration of a gaseous reactant.

- 3) List one way to increase the concentration of a dissolved substance.

- 4) Why does increasing the concentration make a reaction go faster?

- 5) In the reaction where chlorine atoms formed chlorine gas ($2 \text{ Cl} \rightarrow \text{Cl}_2$), why did an increase in temperature cause an increase in the reaction rate?

(over)

- 6) In the reaction between hydrogen and chlorine to produce hydrogen chloride ($\text{H}_2 + \text{Cl}_2 \rightarrow 2 \text{HCl}$), you had to increase the temperature before ANY reaction could occur. This was not true in the previous reaction ($2 \text{Cl} \rightarrow \text{Cl}_2$). Why?