

Binary Ionic with Roman Numerals

Name_____

<u>Name</u>	<u>Ions</u>		<u>Formula</u>
1) silver chloride	_____	_____	_____
2) _____	_____	_____	AlCl ₃
3)_____	Zn ⁺²	Br ⁻¹	_____
4) copper(I)oxide	_____	_____	_____
5) copper(II)oxide	_____	_____	_____
6)_____	_____	_____	FeO
7)_____	_____	_____	Fe ₂ O ₃
8) _____	_____	_____	Li ₂ S
9) Lead(II) flouride	_____	_____	_____
10) _____	_____	_____	PbCl ₄

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1) silver chloride	$\text{Ag}^{+1} \quad \text{Cl}^{-1}$	AgCl
2) Aluminum Chloride	$\text{Al}^{+3} \quad \text{Cl}^{-1}$	AlCl_3
3)Zinc Bromide	$\text{Zn}^{+2} \quad \text{Br}^{-1}$	ZnBr_2
4) copper(I)oxide	$\text{Cu}^{+1} \quad \text{O}^{-2}$	Cu_2O
5) copper(II)oxide	$\text{Cu}^{+2} \quad \text{O}^{-2}$	CuO
6)iron(II)oxide	$\text{Fe}^{+2} \quad \text{O}^{-2}$	FeO
7)Iron(III)oxide	$\text{Fe}^{+3} \quad \text{O}^{-2}$	Fe_2O_3
8) Lithium Sulfide	$\text{Li}^{+1} \quad \text{S}^{-2}$	Li_2S
9) Lead(II) flouride	$\text{Pb}^{+2} \quad \text{F}^{-1}$	PbF_2
10) Lead(IV)chloride	$\text{Pb}^{+4} \quad \text{Cl}^{-1}$	PbCl_4