

Solubility Solutions

Name_____

Below is a set of solutions. Using the solubility rules sheet fill in the table with the results from combining each of the compounds with each other. Fill in the table using following codes:

P = precipitate forms

G = a gas is formed

N= nothing happens

(A piece of information you need to know: If an acid is mixed with a carbonate, carbon dioxide gas is formed.)

These are the compounds:

A = Na_2CO_3

B = BaCl_2

C = AgNO_3

D = HCl - (hydrochloric acid)

E = $\text{Cu}(\text{NO}_3)_2$

	A	B	C	D	E
A					
B					
C					
D					
E					

On the back, write balanced equations for each reaction which resulted in a precipitate.

Given the following completed table, label each of the following compounds with the correct letter.

___ H_2O

___ NaCl

___ BaCl_2

___ HCl (hydrochloric acid)

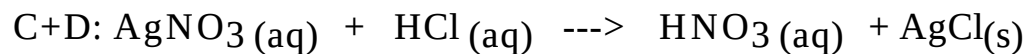
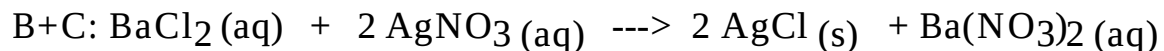
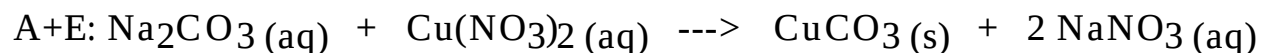
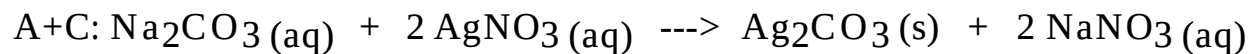
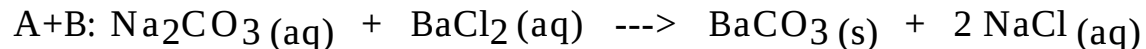
___ Na_2CO_3

___ H_2SO_4 (sulfuric acid)

___ AgNO_3

	A	B	C	D	E	F	G
A		P	P	N	N	N	P
B	P		P	N	G	N	G
C	P	P		N	P	P	P
D	N	N	N		N	N	N
E	N	G	P	N		N	N
F	N	N	P	N	N		N
G	P	G	P	N	N	N	

Solubility Solutions Answers



D H₂O

F NaCl

A BaCl₂

E HCl (hydrochloric acid)

B Na₂CO₃

G H₂SO₄ (sulfuric acid)

C AgNO₃