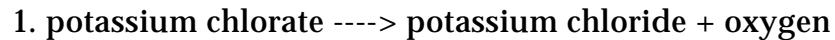


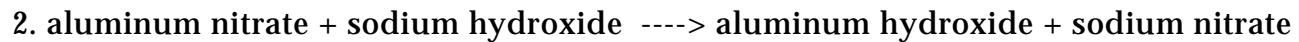
TYPES OF REACTIONS

NAME _____

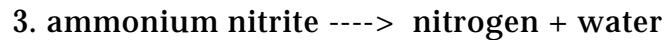
Write and balance the equations for each reaction as they are assigned. Also tell what type of reaction each one is.



Type _____



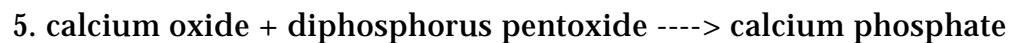
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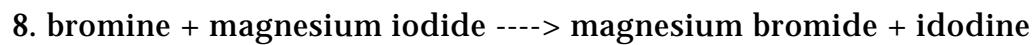
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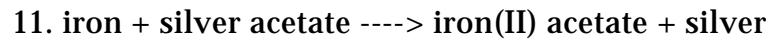
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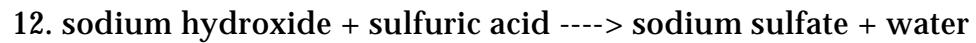
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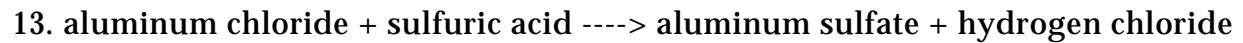
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Types of Reactions Answer Key

1. potassium chlorate ----> potassium chloride + oxygen
 - $2\text{KClO}_3 \longrightarrow 2\text{KCl} + 3\text{O}_2$
 - Decomposition
2. aluminum nitrate + sodium hydroxide ----> aluminum hydroxide + sodium nitrate
 - $\text{Al}(\text{NO}_3)_3 + 3\text{NaOH} \longrightarrow \text{Al}(\text{OH})_3 + 3\text{NaNO}_3$
 - Double Replacement
3. ammonium nitrite ----> nitrogen + water
 - $\text{NH}_4\text{NO}_2 \longrightarrow \text{N}_2 + 2\text{H}_2\text{O}$
 - Decomposition
4. iron (III) bromide + ammonium sulfide ----> iron (III) sulfide + ammonium bromide
 - $2\text{FeBr}_3 + 3(\text{NH}_4)_2\text{S} \longrightarrow \text{Fe}_2\text{S}_3 + 6\text{NH}_4\text{Br}$
 - Double Replacement
5. calcium oxide + diphosphorus pentoxide ----> calcium phosphate
 - $3\text{CaO} + \text{P}_2\text{O}_5 \longrightarrow \text{Ca}_3(\text{PO}_4)_2$
 - Synthesis
6. aluminum + copper (II) chloride ----> aluminum chloride + copper
 - $2\text{Al} + 3\text{CuCl}_2 \longrightarrow 2\text{AlCl}_3 + 3\text{Cu}$
 - Single Replacement
7. calcium hydroxide + nitric acid ----> calcium nitrate + water
 - $\text{Ca}(\text{OH})_2 + 2\text{HNO}_3 \longrightarrow \text{Ca}(\text{NO}_3)_2 + 2\text{H}_2\text{O}$
 - Double Replacement
8. bromine + magnesium iodide ----> magnesium bromide + iodine
 - $\text{Br}_2 + \text{MgI}_2 \longrightarrow \text{MgBr}_2 + \text{I}_2$
 - Single Replacement
9. sodium bicarbonate ----> sodium oxide + carbon dioxide + water
 - $2\text{NaHCO}_3 \longrightarrow \text{Na}_2\text{O} + 2\text{CO}_2 + \text{H}_2\text{O}$
 - Decomposition
10. aluminum + oxygen ----> aluminum oxide
 - $4\text{Al} + 3\text{O}_2 \longrightarrow 2\text{Al}_2\text{O}_3$
 - Synthesis
11. iron + silver acetate ----> iron (II) acetate + silver
 - $\text{Fe} + 2\text{AgC}_2\text{H}_3\text{O}_2 \longrightarrow \text{Fe}(\text{C}_2\text{H}_3\text{O}_2)_2 + 2\text{Ag}$
 - Single Replacement
12. sodium hydroxide + sulfuric acid ----> sodium sulfate + water
 - $2\text{NaOH} + \text{H}_2\text{SO}_4 \longrightarrow \text{Na}_2\text{SO}_4 + 2\text{H}_2\text{O}$
 - Double Replacement
13. aluminum chloride + sulfuric acid ----> aluminum sulfate + hydrogen chloride
 - $2\text{AlCl}_3 + 3\text{H}_2\text{SO}_4 \longrightarrow \text{Al}_2(\text{SO}_4)_3 + 6\text{HCl}$
 - Double Replacement